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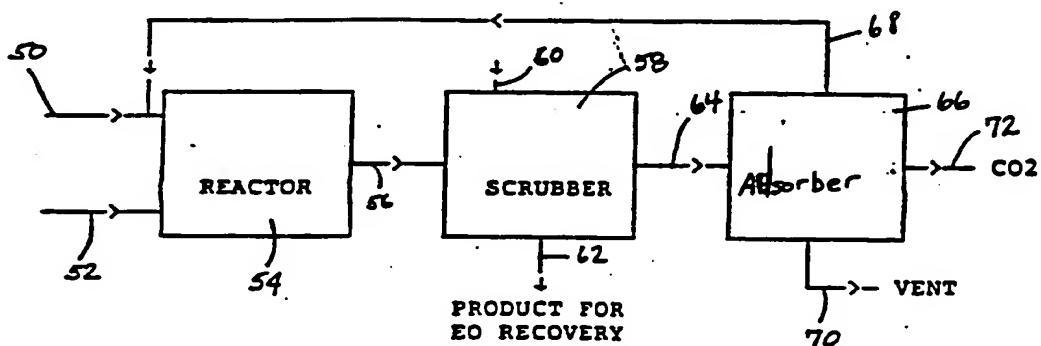
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(54) Process for the production of alkylene oxide.

(57) In a process for the production of an alkylene oxide, an alkene and an oxygen containing gas (oxygen, air or oxygen-enriched air) are reacted in a reactor 54 in the presence of a flame suppressor under conditions of low alkene conversion and high alkylene oxide selectivity. Alkylene oxide is separated from the gas mixture in a scrubber 58. The gases that pass through the scrubber 58 are separated in a pressure swing adsorber 66 into a first stream enriched in unreacted alkene and a second stream depleted of alkene. A third stream comprising carbon dioxide may be produced. The first stream is recycled to the reactor 54.

FIG 3.,



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The present invention is directed to a process for the production of alkylene oxides from alkenes and an oxygen-containing gas in which unreacted alkenes are recovered and recycled to improve the process efficiency and the off-gases are treated without incineration which saves natural resources and provides for the accumulation of carbon dioxide as a by-product.

5 The production of alkylene oxides from alkenes is well known. Brian J. Ozero, Handbook of Chemicals Production Processes, edited by Robert Meyers, McGraw Hill Book Co. (1986) at Chapter 1.5, discloses cyclic processes using both oxygen and air as an oxidant for the production of ethylene oxide from ethylene. In these processes, the alkene is oxidised in a multitubular catalytic reactor in vapour phase. The reactor off gases are cooled and scrubbed with water in an absorber
10 to recover ethylene oxide which is sent to a recovery section for further purification.

In the oxygen-based process described by Ozero, the scrubber off gases are divided into three parts which are respectively: (i) recycled to the reactor, (ii) vented and (iii) sent to a separator for carbon dioxide removal and recycle of the remaining hydrocarbons. This process suffers from several disadvantages. In particular, the process requires a separate carbon dioxide removal unit and a purge to remove argon which
15 would otherwise accumulate in the system.

In the air-based process described by Ozero, the scrubber off gases are sent to a second reactor, which is the purge reactor, where additional unreacted ethylene is reacted using a higher air to ethylene ratio, thereby foregoing some ethylene oxide selectivity. The reactor off gases are passed through another water scrubber to recover ethylene oxide.
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It is known that the volume of hydrocarbons purged, when utilising air as a source of oxygen, requires that the purge scrubber off gases be incinerated to remove any remaining hydrocarbons in order to meet environmental regulations. In this air-based process, an additional purge oxidation reactor, a water scrubber, and an effluent incinerator are required, as well as a greater volume of catalyst. Another shortcoming of the processes described by Ozero is that they are for practical purposes limited to the use of either oxygen or air. It would be advantageous to eliminate the purge and additional carbon dioxide separator and operate
25 the ethylene oxide reactor at a higher selectivity to improve the overall process efficiency.

According to the present invention there is provided a process for the production of an alkylene oxide from a corresponding alkene comprising the steps of:

- a) feeding the alkene, an oxygen-containing gas and a flame suppressor to a reaction zone;
- 30 b) reacting the alkene and oxygen containing gas in said reaction zone under conditions of low alkene conversion and high alkylene oxide selectivity to produce a mixture of the alkylene oxide and off gases;
- c) quenching the mixture obtained in step (b) and removing substantially all of the alkylene oxide from said quenched mixture;
- d) separating the off-gases in said quenched mixture into a first stream enriched in unreacted alkene and
35 a second stream lean in unreacted alkene, and
- e) forwarding the first stream to the reaction zone and removing the second stream.

That the process of the present invention employs a separation system in which substantially all of the unreacted alkene is removed from the scrubber off gases and recycled back to the reactor enables the reaction to be conducted at low alkene conversion, high alkylene oxide selectivity while the separation
40 system off gases may be vented directly. The flame suppressor can be fed either continuously or only during start-up.

As used herein, the terms "low conversion" and "high selectivity" means rates of conversion and selectivity which are respectively lower and higher than processes typically practised in the prior art. More specifically, the term "low conversion" is a rate of conversion which results in an increase in the rate of
45 selectivity of at least 1% compared to conventional once-through processes which do not recycle the alkene to the oxidation reactor. By way of example only, the rates typically employed in the present invention for the conversion of alkene to alkylene oxide are in the range of from about 5 to 80% most typically from about 10 to 60%. Corresponding selectivity rates are in the range from about 50 to 90% for ethylene oxide and 15 to 80% for propylene oxide.

50 The mixture is quenched to remove the alkylene oxide produced for further purification in a manner known to those skilled in the art.

The remaining off gases are sent to at least one separator, either with or without compression, to produce a first stream containing unreacted alkene and the flame suppressor and a second stream containing the remaining gases. The first stream is recycled back to the reactor to be combined with fresh
55 oxygen-containing gas and fed into the reactor. The second stream is removed and may be vented, incinerated or further separated to remove purified carbon dioxide.

In accordance with the present invention oxygen-enriched air can be used as the oxidant because the system provides for the effective removal of nitrogen gas while taking advantage of the positive effects of

conducting the process under high alkylene oxide selectivity at low alkene conversion rates.

In addition, the efficient removal of nitrogen and carbon dioxide from the system reduces the outlay for capital equipment by making the flow rates lower and thereby eliminating the need for a separate carbon dioxide removal system.

5 Further, when oxygen enriched air is used as the oxidant in the present invention, it is not necessary to use methane to render the reactants inflammable as customarily employed in prior art systems based on oxygen. This is because the nitrogen present in the air and the recycled carbon dioxide provide sufficient flame suppression without undesirable nitrogen gas build-up.

10 The known processes will now be further described with reference to Figures 1 and 2 of the accompanying drawings. Processes according to the invention are then described by way of example with reference to Figures 3 and 4 of the accompanying drawings. Like reference characters indicate like parts in the drawings. Processes described in connection with the production of ethylene oxide from ethylene are applicable to the production of propylene oxide from propylene and the like.

In the drawings:

15 FIGURE 1 is a schematic view of a prior art system for converting ethylene to ethylene oxide using air as the oxidant;

FIGURE 2 is a schematic view of a prior art system for converting ethylene to ethylene oxide using pure oxygen as the oxidant;

20 FIGURE 3 is a schematic view of an embodiment of the present invention for converting an alkene to the corresponding alkylene oxide using pure oxygen or other oxygen-containing gas as the oxidant; and

25 FIGURE 4 is a schematic view of another embodiment of the present invention similar to the embodiment shown in FIGURE 3 in which plural adsorbers are used for separation.

Prior processes for the production of ethylene oxide by the oxidation of ethylene have employed either air or pure oxygen as the oxidant. FIGURE 1 shows a prior art system using air as the oxidant and a purge 25 to remove the inert gases to prevent nitrogen build up and an incinerator to meet environmental regulations.

Specifically air and ethylene are fed to an oxidation reactor 2 containing a catalyst composed, for example, of a metal on a support such as silver on alumina. A mixture of ethylene oxide and off gases are cooled and then fed to a scrubber 4 in which water is used to dissolve the ethylene oxide for subsequent treatment. The off gases including unreacted ethylene are removed and divided into two streams. A first stream is returned via a line 6 to the oxidation reactor 2 while a second stream is sent via a line 8 to a purge reactor 10.

The second stream containing off gases including oxygen, nitrogen, argon, ethylene and carbon dioxide is combined with additional quantities of air in the purge reactor 10 to provide a relatively high oxygen to ethylene ratio which obtains a higher conversion of ethylene and thereby produces additional quantities of ethylene oxide. The by-products of the reaction are sent to a second scrubber 12 in which ethylene oxide is recovered and a stream of off gases containing unreacted ethylene is divided into two streams.

A first stream is fed via a line 14 back to the purge reactor 10 and the second stream is sent to an incinerator 16, where the off-gases, particularly the hydrocarbons contained therein, are combusted and thereafter vented.

40 Referring to FIGURE 2, there is shown a prior art system in which pure oxygen gas is used as the oxidant. Ethylene, oxygen and a flame suppressor such as methane gas are sent to the oxidation reactor 2 of the same type described in connection with FIGURE 1. Ethylene oxide and off gases are sent to the scrubber 4 for recovery of ethylene oxide for purification. The off-gases are divided into three streams, one stream flows via the line 6 back to the oxidation reactor 2. A second stream is sent to an incinerator for 45 combusting the hydrocarbons and a third stream is sent to an absorber 18 for removing carbon dioxide from the off gases. A portion of the off gases removed from the absorber 18 is sent to the oxidation reactor 2 and the remaining off gases are recycled to the absorber 18.

In accordance with the present invention, there is provided a system for the conversion of an alkene to an alkylene oxide in which the oxidant can be selected from any one or more of pure oxygen, air and oxygen enriched air. The oxidants can be combined and the composition of the oxidants can be changed without the need for material changes in capital equipment depending on the alkylene oxide requirement. As a consequence, the system of the present invention provides greater flexibility in the use of oxidants over known systems.

Referring to FIGURE 3, the process of the present invention commences by forwarding a gaseous 55 alkene via a line 50 and an oxygen containing gas via a line 52 to an oxidation reactor 54. The starting alkenes have from 2 to 4 carbon atoms, particularly ethylene and propylene.

The oxidation reactor contains a suitable oxidation catalyst, such as silver on alumina, in a fixed, fluidized or slurry reactor. The catalyst may be promoted with other known metals to improve stability and

selectivity.

As previously indicated, the oxidant may range from pure oxygen to air. The optimum oxygen concentration will depend on whether the process is used to retrofit an existing plant or implemented as a new plant or, if retrofitted, the need for additional capacity. In other words, the process can be employed without major modifications in plants of varying capacity.

It is also necessary in accordance with the present invention to provide for a flame suppressor. In the case of pure oxygen, methane or ethane may be used as a flame suppressor. The amount of the flame suppressor is controlled so as to avoid the formation of flammable mixtures in the system. Typically, the total amount of flame suppressors is in the range of about 20 to 80%. A major portion of the flame suppressor is added only during the start up since most of the flame suppressor is recycled. For the process using air as the oxygen containing gas, typical nitrogen concentrations are about 30% by volume and carbon dioxide concentrations about 20% as the reactor feed.

The oxidation reaction is conducted at a temperature in the range of from about 200 to 500C and a pressure of from about 15 to 400 psig.

The resulting product mixture includes the alkylene oxide (e.g. ethylene oxide) unreacted alkene (e.g. ethylene), oxygen and carbon dioxide, nitrogen, and argon if other than pure oxygen is used as the oxidant.

The mixture is cooled in a cooler (not shown) and then sent via the line 56 to a scrubber 58 wherein water from a line 60 is used to separate the alkylene oxide from the off gases. The alkylene oxide is removed from the scrubber 58 through a line 62.

The off gases are sent via a line 64 to a pressure swing adsorber 66 containing two or more beds, preferably in parallel with suitable adsorbents capable of removing carbon dioxide, nitrogen and argon, if present in the reactor feed. Typical adsorbents include activated carbon, silica gel and molecular sieves and other adsorbents well known to those skilled in the art. The scrubber off gases enter the adsorber 66 at a temperature of from about 10 to 100C and a pressure of from about 0 to 400 psig. Depending upon the reactor pressure, it may be necessary to compress the scrubber off gases before feeding it to the pressure swing adsorber 66.

A first stream exits the adsorber 66 via a line 68 for return to the oxidation reactor 54. The first stream contains substantially all of the unreacted alkylene and minor amounts of carbon dioxide, nitrogen, oxygen and argon.

Because substantially all of the hydrocarbons (e.g. ethylene) leaving the adsorber 66 are returned in the recycle, a second stream containing off gases excluding hydrocarbons can be vented via the line 70 without incineration. As a consequence, the process of the present invention can operate without the costly incineration apparatus associated with prior art processes.

Carbon dioxide can be removed as part of the vent gases or separated from the vent gases and removed as a by-product via a line 72 depending upon the oxygen concentration in the reactor feed. If air is used as the oxygen containing gas, carbon dioxide may not be recovered as a by-product. The separation of carbon dioxide from the vent gases requires a carbon dioxide adsorbing material such as molecular sieve. Carbon dioxide is removed by adsorbing it on the molecular sieve preferentially over the remaining gases and is obtained as a desorbed product.

When air is used as the oxygen containing gas or when very large quantities of carbon dioxide must be removed, it may be desirable to use more than one pressure swing adsorption system in series. The first system preferably is capable of selectively adsorbing carbon dioxide while the second preferentially adsorbs hydrocarbons.

Referring to FIGURE 4, there is shown the use of two pressure swing adsorption columns. The off gases from the scrubber 58 are sent via the line 64 to a first pressure swing adsorber 80 containing adsorbents which preferentially adsorb carbon dioxide as described previously. Carbon dioxide is removed as a by-product via a line 82. The remaining off gases are sent via a line 84 to a second pressure swing adsorber 86 containing adsorbents which preferentially remove hydrocarbons (e.g. ethylene) from the off gas stream via a line 88 for recycling. Nitrogen, oxygen and other off gases can be vented from the second pressure swing adsorber 86 via a line 90 without incineration.

EXAMPLE 1

The process of the present invention was conducted in accordance with FIGURE 3 in the following manner to produce ethylene oxide. 141 moles of ethylene, and 1203.5 moles of air (containing 252.6 moles of oxygen, 950.2 moles of nitrogen and a trace amount of ethylene) were forwarded via the lines 50, 52 respectively into the oxidation reactor 54. In addition, the reactor 54 was supplied with a recycle of ethylene and other off gases via the line 68 to raise the quantity of the gases therein to that shown in Table 1.

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TABLE 1

CONTENTS TO THE REACTOR		
Gas	Moles	% by volume
ethylene	1293.7	35.7
ethane	72.4	2.0
oxygen	281.2	7.8
carbon dioxide	792.9	21.9
nitrogen	1187.8	32.7

The gas mixture set forth in Table 1 produces gas mixture shown in Table 2 as the product. This product was forwarded via the line 56 to the scrubber 58.

TABLE 2

CONTENTS TO THE SCRUBBER		
Gas	Moles	% by volume
ethylene	1164.3	32.5
ethane	72.4	2.0
oxygen	143.1	4.0
ethylene oxide	100.0	2.8
carbon dioxide	851.6	23.8
water vapour	58.7	1.6
nitrogen	1187.8	33.2

100 moles of ethylene oxide were removed from the scrubber 58 via the line 62 to provide a conversion rate of ethylene to ethylene oxide of 10.0% and a selectivity of 77%.

After quenching, the gases were sent to a pressure swing adsorber 66 to separate ethylene and, optionally carbon dioxide from the off gases. The charge sent to the pressure swing adsorber had the composition shown below in Table 3.

TABLE 3

CONTENTS TO THE PSA		
Gas	Moles	% by volume
ethylene	1164.3	33.8
ethane	72.4	2.1
oxygen	43.1	4.2
carbon dioxide	881.0	25.5
nitrogen	1187.8	34.4

The temperature in the pressure swing adsorber 66 was in the range from about 15 to 35C and a pressure of from about 5 to 100 psig.

Substantially all of the ethylene (1152.7 moles; 99 + %) was sent via the line 68 to the oxidation reactor 54. A gas mixture containing 11.6 moles of ethylene, 0.7 moles of ethane, 74.9 moles of oxygen and 88.1 moles of carbon dioxide was vented out of the system via the line 70. Depending upon the hydrocarbon recovery in the pressure swing adsorber, it may be necessary to incinerate the vent stream.

EXAMPLE 2

The process of the present invention was conducted in accordance with FIGURE 3 using pure oxygen as the oxidant to produce ethylene oxide. 141.0 moles of ethylene, 213.0 moles of oxygen and a trace

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amount of ethane was forwarded via the lines 50 and 52, respectively to the oxidation reactor 54. In addition, the reactor 54 was supplied with a recycle of ethylene and other gases via the line 68 to raise the quantity of gases therein to that shown in Table 4.

TABLE 4

CONTENTS TO THE REACTOR		
Gas	Moles	% by volume
ethylene	1293.7	54.1
ethane	72.4	3.0
oxygen	231.7	9.7
carbon dioxide	792.9	33.2

The gas mixture shown in Table 4 was reacted to produce the stream shown in Table 5 which was forwarded via the line 56 to the scrubber 58.

TABLE 5

CONTENTS TO THE SCRUBBER		
Gas	Moles	% by volume
ethylene	1164.3	49.7
ethane	72.4	3.1
oxygen	93.6	4.0
ethylene oxide	100.0	4.3
carbon dioxide	851.6	36.4

The ethylene oxide was removed from the scrubber 58 to provide a conversion rate of ethylene to ethylene oxide of 10.0% and a selectivity of 77%.

After quenching, the gases were sent to a pressure swing adsorber 66 to separate ethylene and, optionally carbon dioxide from the off gases. The charge sent to the pressure swing adsorber had the composition shown below in Table 6.

TABLE 6

CONTENTS TO THE PSA		
Gas	Moles	% by volume
ethylene	1164.3	52.7
ethane	72.4	3.3
oxygen	93.6	4.2
carbon dioxide	881.0	39.8

The temperature in the adsorber 66 was the range from about 15 to 35 °C and a pressure of from about 5 to 100 psig.

Substantially all of the ethylene (1152.7 moles; 99+ % by volume) was sent via the line 68 to the oxidation reactor 54. A gas mixture containing 11.6 moles of ethylene, 0.7 moles of ethane and 74.9 moles of oxygen and 88.1 moles of carbon dioxide was vented out of the system via the line 70. Depending on the amount of hydrocarbon in the stream, it may be necessary to incinerate the vent stream.

EXAMPLE III

The process of the present invention was conducted in accordance with FIGURE 3 in the following manner to produce propylene oxide. 405.1 moles of propylene, 8.3 moles of propane and 558.4 moles of

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oxygen, were forwarded via the lines 50, 52 respectively into the oxidation reactor 54. In addition, the reactor 54 was supplied with a recycle of propylene and other off gases via the line 68 to raise the quantity of the gases therein to that shown in Table 7.

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TABLE 7

CONTENTS TO THE REACTOR		
Gas	Moles	% by volume
propylene	2461.5	58.5
propane	698.9	16.6
oxygen	625.7	14.9
carbon dioxide	336.2	8.0
ethylene	84.9	2.0
formaldehyde	1.9	0.0

The gas mixture set forth in Table 7 produces the gas mixture shown in Table 8 as the product. This product was forwarded via the line 56 to the scrubber 58.

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TABLE 8
CONTENTS TO THE SCRUBBER

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<u>Gas</u>	<u>Moles</u>	<u>% by volume</u>
propylene	2082.5	49.7
propane	698.9	16.7
oxygen	83.8	2.0
propylene oxide	100.0	2.4
acetaldehyde	134.6	3.2
formaldehyde	64.9	1.5
carbon dioxide	594.8	14.2
ethylene	153.8	3.7
water vapour	241.0	5.7
Balance (alcohol, acetone, dienes)	38.0	0.9

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100 moles of propylene oxide were removed from the scrubber 58 via the line 62 to provide a conversion rate of propylene to propylene oxide of 15.4% and a selectivity of 26.4%.

After quenching, the gases were sent to a pressure swing adsorber 66 to separate propylene and, optionally carbon dioxide from the off gases. The charge sent to the pressure swing adsorber had the composition shown below in Table 9.

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TABLE 9

CONTENTS TO THE PSA		
Gas	Moles	% by volume
propylene	1300.6	52.0
propane	413.8	16.5
oxygen	82.8	3.3
carbon dioxide	560.3	22.4
ethylene	141.5	5.7
formaldehyde	3.2	0.1

The temperature in the pressure swing adsorber 66 was in the range from about 15 to 35C and a
 15 pressure of from about 5 to 100 psig.

1274.6 moles of propylene was sent via the line 68 to the oxidation reactor 54. A gas mixture containing
 20 26.0 moles of propylene, 8.3 moles of propane, 16.6 moles of oxygen, 224.1 moles of carbon dioxide and
 25 56.6 moles of ethylene was vented out of the system via the line 70. Depending upon the hydrocarbon
 recovery in the pressure swing adsorber, it may be necessary to incinerate the vent stream.

Claims

1. A process for the production of an alkylene oxide from a corresponding alkene comprising the steps of:
 - a) feeding the alkene, an oxygen containing gas and a flame suppressor to a reaction zone;
 - b) reacting the alkene and oxygen containing gas in said reaction zone under conditions of low alkene conversion and high alkylene oxide selectivity to produce a mixture of the alkylene oxide and off gases;
 - c) quenching the mixture obtained in step (b) and removing substantially all of the alkylene oxide from said quenched mixture;
 - d) separating the off-gases in said quenched mixture into a first stream enriched in unreacted alkene and a second stream lean in unreacted alkene, and
 - e) forwarding the first stream to the reaction zone and removing the second stream.
2. A process according to Claim 1, further comprising venting or incinerating the second stream.
3. A process according to Claim 1 or Claim 2, further comprising separating the off gases into a third stream containing carbon dioxide.
4. A process according to any one of the preceding claims, wherein step (d) comprises forwarding the quenched mixture to a pressure swing adsorber containing an adsorbent capable of preferentially adsorbing hydrocarbons and selectively desorbing hydrocarbons therein.
5. A process according to Claim 4, wherein step (d) comprises forwarding the quenched mixture to a pressure swing adsorber containing a carbon dioxide adsorbent and selectively desorbing carbon dioxide therein.
6. A process according to Claim 5, further comprising removing carbon dioxide as a by-product from the pressure swing adsorber.
7. A process according to any one of claims 1 to 3, wherein step (d) comprises forwarding the quenched mixture to a first pressure swing adsorber containing a carbon dioxide adsorbent or a hydrocarbon adsorbent, forwarding said off gas stream to a second pressure swing adsorber containing the other of said carbon dioxide adsorbent and hydrocarbon adsorbent; and separating and recycling the hydrocarbon to the reaction zone.
8. A process according to Claim 7, wherein an off gas stream comprising nitrogen and oxygen is formed in the second pressure swing adsorber, and further comprising venting said off gas stream.

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9. A process according to any one of the preceding claims, wherein the oxygen containing gas is air and the flame suppressor is a combination of nitrogen gas present in the air and carbon dioxide present in the first stream.
- 5 10. A process according to any one of claims 1 to 8, wherein the oxygen containing gas is oxygen gas or oxygen enriched air and the flame suppressor is an alkane alone or in combination with nitrogen gas.
11. A process according to any one of the preceding claims, wherein the total amount of flame suppressor is in the range of about 20 to 80% by volume.
- 10 12. A process according to any one of the preceding claims wherein the alkene is ethylene or propylene.
13. A process according to claim 7, wherein the carbon dioxide adsorbent is a carbon molecular sieve and the hydrocarbon adsorbent is a molecular sieve, activated carbon or silica gel.
14. A process according to any one of the preceding claims, wherein the reaction of the alkene and the oxygen containing gas is conducted at a temperature of from about 200 to 500 °C and a pressure of from about 15 to 400 psig, and said off-gases are separated at a temperature of from about 10 to 100 °C and a pressure of from about 0 to 400 psig.

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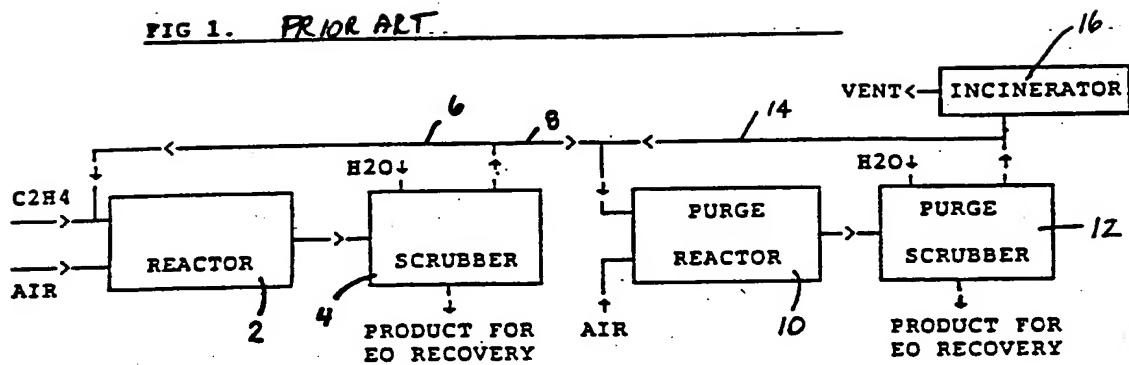
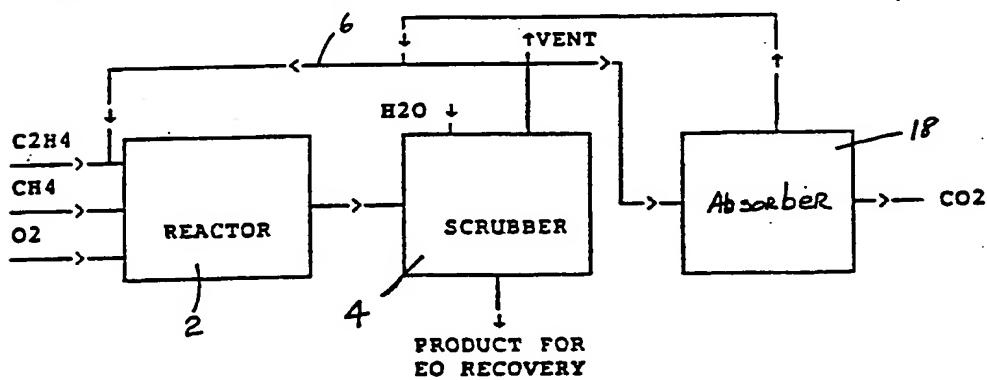
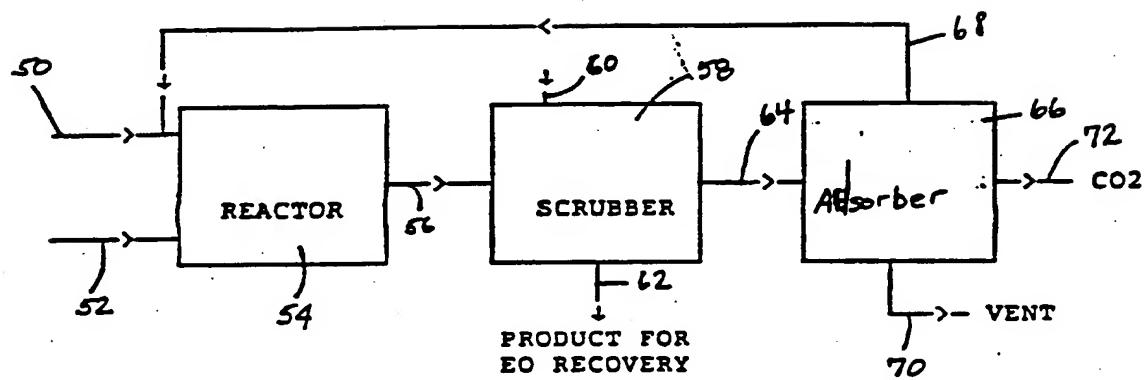
FIG 1. PRIOR ARTFIG 2. PRIOR ART

FIG 3.



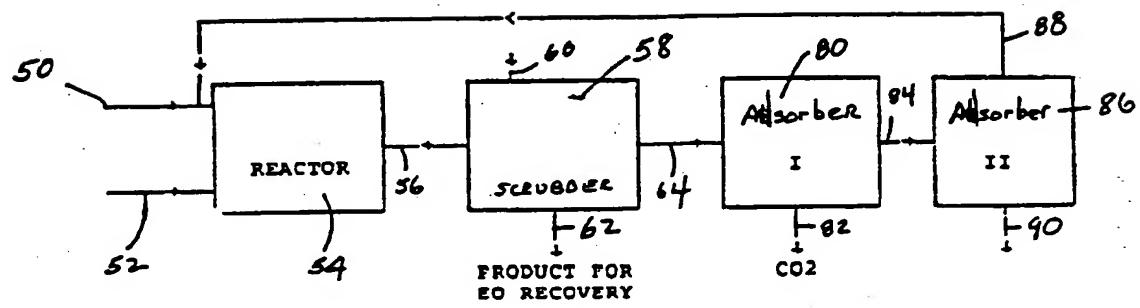


Fig. 4



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REPORT

Application Number

EP 91 30 5662

DOCUMENTS CONSIDERED TO BE RELEVANT					
Category	Citation of document with indication, where appropriate, of relevant passages	Relevant to claim	CLASSIFICATION OF THE APPLICATION (Int. Cl.5)		
X	EP-A-0 328 280 (THE BOC GROUP, INC.) * the whole document, particularly page 6, lines 13-16 and page 11, claim 16 ** -----	1-14	C 07 D 301/10 C 07 D 301/32		
X	US-A-2 241 019 (F.J. METZGER) * the whole document ** -----	1-14			
X	US-A-4 879 396 (B.J. OZERO) * the whole document ** -----	1-14			
X	US-A-3 266 221 (W.F. AVERY) * the whole document ** -----	1-14			
X	US-A-3 867 113 (E.G. FORSTER ET AL.) * the whole document ** -----	1-14			
The present search report has been drawn up for all claims					
Place of search	Date of completion of search	Examiner			
The Hague	26 November 91	ALLARD M.S.			
CATEGORY OF CITED DOCUMENTS					
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